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State the Kelvin-Planck and Clausius statements of second law of thermodynamics and prove their equivalence. 6. 0.5 kg of air at 1 bar and 47°C is compressed in a piston-cylinder assembly to 4 bar and 127°C by doing 5 kJ of work when the surrounding temperature

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Answer: Heat flows into the system, q is +ve. Heat flows out of the system, q is -ve. Work is done on the system, w is +ve. Work is done by the system, w is -ve. Q34: Define First law of thermodynamics. Answer: The total energy of an isolated system remains constant though it may change from one form to another.

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Solution: $N_2(g) + 3H_2 \rightarrow 2NH_3(g)$ $\Delta S = 2 - 4 = -2 < 0$. So entropy is negative. Hence option (c) is the answer. 3. If an endothermic reaction is non-spontaneous at freezing point of water and becomes feasible at its boiling point, then. (a) ΔH is -ve, ΔS is +ve. (b) ΔH and ΔS both are +ve.

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