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Chemical reaction engineering is concerned with the exploitation of chemical reactions on a commercial scale. Its goal is the successful design and operation of chemical reactors. This text emphasizes qualitative arguments, simple design methods, graphical procedures, and frequent comparison of capabilities of the major reactor types.

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Octave Levenspiel was a professor of the field Chemical engineering at Oregon State University. In this vast and evergreen field his major interests lied in Chemical Reaction Engineering which is one of the core subjects in Chemical Engineering. He published this book which is considered as a Bible for understanding major concepts of Chemical Reaction Engineering.

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1 Chemical reactions 1.1 Rate of reaction and dependence on temperature We will once again look at the formation of ammonia (NH<sub>3</sub>) from nitrogen and hydrogen (see section Chemical equilibrium of the thermodynamics chapter). This reaction follows the equation: N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub> (1) H<sub>0</sub> = 92 kJ mol<sup>-1</sup> S<sub>0</sub> = 192 J mol<sup>-1</sup> K To nd the Gibbs free energy of ...

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By sizing a chemical reactor we mean we're either determining the reactor volume to achieve a given conversion or determine the conversion that can be achieved in a given reactor type and size. Here we will assume that we will be given  $-r_A = f(X)$  and  $F_{A0}$ . In chapter 3 we show how to find  $-r_A = f(X)$ . Given  $-r_A$  as a function of conversion,  $-r_A = f(X)$ , one can size any type of reactor.

Elements of Chemical Reaction Engineering

Octave Levenspiel was a professor of chemical engineering at Oregon State University. His principal interest was chemical reaction engineering, and he was the author of a major textbook Chemical Reaction Engineering as well as numerous research publications.

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