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expression importance of the
magnitude of K how K depends on rxn
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Systems. Chapter 15 Water and Aqueous Systems. The Water Molecule: a Review. Water is a simple tri-atomic molecule, H₂O. Each O-H bond is highly polar, because of the high electronegativity of the oxygen (N, O, F, and Cl have high values) bond angle of water =

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compounds are electrolytes because they dissociate into ions and conduct an electrical current in aqueous solution or in the molten state.

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System. the chaotic movement of colloidal particles, caused by collision with particles of the solvent in which they are dispersed. a compound that conducts an electric current when it is in an aqueous solution or in the molten state; all ionic compounds are electrolytes, but most covalent

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chapter. heterogeneous mixture containing particles that are small enough to remain dispersed in the solvent and do not separate on standing.

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Chemistry, Chapter 15, Water and
Aqueous Systems. surface tension.
surfactant. aqueous solution. solvent.
the inward force or pull that tends to
minimize the surface area. any
substance that interferes with

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hydrogen bonding between water molecules is
water that contains dissolved
substances. the dissolving medium in
a solution.

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Chapter 15 Water and Aqueous
Systems 159 SECTION 15.1 WATER

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AND ITS PROPERTIES (pages
445-449) This section describes the
properties of water in the liquid and
solid states and explains how
hydrogen bonding affects the surface
tension and vapor pressure of water.
Water in the Liquid State (pages
445-447) 1.

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SECTION 15.1 WATER AND ITS
PROPERTIES (pages 445–449)
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source:eur-lex.europa.eu He may want
to stretch himself once a worker knows

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his efforts do not go unnoticed. For instance, if he knows his performance will be judged based on achievement of a target, he will work harder to achieve it.

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crystal of solid NaCl is placed into an
aqueous NaCl solution. No precipitate
forms in the bottom of the container.
The final solution is: a.) super
saturated ... Chapter 5: Soil Water 57

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2. Section 15.1 Water and its Properties

OBJECTIVES: Explain the high surface tension and low vapor pressure of water in terms of the structure of the water molecule and hydrogen bonding.

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enough to remain dispersed in the
solvent and do not separate on
standing.

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CHAPTER 15 | Aqueous Equilibria:
Chemistry of the Water World 15.1.

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P15 shows .1 four lines to describe the possible dependence of percent ionization of acetic acid with concentration. We are to choose the one that best represents the trend for this weak acid. Analyze

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