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Chapter 11 Chemical Reactions- Chemistry by Ms.Basima

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CHAPTER 11 SECTIONS 1 Defining Stoichiometry 2 Stoichiometric Calculations 3 Limiting Reactants 4 Percent Yield LaunchLAB What evidence can you observe that a reaction has stopped? During a chemical reaction, reactants are consumed as new products form. In this lab, you will look for signs a chemical reaction has stopped. Steps in Stoichiometric Calculations

### CHAPTER 11 Stoichiometry

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Write correct formulas for the products in these double replacement reactions.  $\text{H}_3\text{PO}_4 + 2\text{K}_2\text{CO}_3 + \text{BaCl}_2 \rightarrow$  3)  $\text{Al} + \text{Fe}(\text{NO}_3)_3 \rightarrow$  5)  $\text{AgNO}_3 + \text{KCl} \rightarrow$   $\text{H}_2\text{SO}_4 +$   $\text{Co}(\text{OH})_2 + \text{Pb}(\text{NO}_3)_2 \rightarrow$   $\text{CaS} + \text{Ag}_2\text{CO}_3 + \text{K}_2\text{CrO}_4$

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Chapter 11 • Stoichiometry 367 Start-Up Activities LAUNCH CH 11 Lab What evidence can you observe that a reaction is taking place? During a chemical reaction, reactants are consumed as new products are formed. Often, there are several telltale signs that a chemical reaction is taking place. Procedure 1.

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As shown in Figure 11.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the ammonium dichromate volcano. Heat, light, and gas are produced as a large pile of fluffy green chromium(III) oxide forms. We can describe this reaction with a chemical equation An expression that gives the identities and quantities of the ...

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